

Periodic Table worksheet

Name _____

1. Define a family. _____
2. What is a period? _____
3. What is the symbol for the following elements.
 - a. Magnesium _____
 - b. Potassium _____
 - c. Iron _____
 - d. Copper _____
4. What are the names of the following elements.
 - a. C _____
 - b. Cl _____
 - c. Au _____
 - d. Sr _____
5. What period are the following elements in?
 - a. He _____
 - b. Ge _____
 - c. Rb _____
 - d. I _____
6. What group are the following elements?
 - a. Sulfur _____
 - b. Ca _____
 - c. Iodine _____
 - d. Fe _____
7. Give me an atom with the following characteristics.
 - a. Halogen _____
 - b. Nonmetal _____
 - c. Alkali metal _____
 - d. metalloid _____
 - e. Lanthanide series _____
 - f. Alkaline Earth metal _____
 - g. Transition metal _____
 - h. Nobel gas _____
8. Write the electron configuration for
 - a. Li _____
 - b. Na _____
 - c. K _____

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Use the following clues to determine the elements being described and fill in the missing information.

1. Each atom of this element has 9 protons and 10 neutrons. The atomic number of the element is (1). The mass number is (2). The number of valence electrons is (3). This element can be found in the (4) family on the periodic table. The valence configuration of this atom is (5). This element is classified as a(n) (6). This element is (7), with a symbol of (8).

2. This element is a metalloid of Group IV. It has (9) valence electrons on the third energy level. Its name is (10), symbol (11). Its atomic number is (12) and its mass number is (13). This element has (14) neutrons.

3. This is a Halogen and is found in Group (15). It has (16) valence electrons. It has 17 electrons in a neutral atom. The element is (17), symbol (18). It has (19) protons and a mass number of 35. How many neutrons does it have? (20)

4. This element is the largest of period three. It is (21), symbol (22). Its atomic number is (23) and it is found in the family (24).

5. Each atom of this element has 56 protons and 81 neutrons. The atomic number is (25) and the atomic mass number is (26). The number of valence electrons this atom has is (27) because it is found in Group (28). The element is (29), symbol (30).

- 1. _____
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- 27. _____
- 28. _____
- 29. _____
- 30. _____

6. This is a Group II element with 12 protons. The number of electrons in a neutral atom of this element is (31). It has (32) valence electrons on the (33) energy level. The family name of this element is (34). The mass number is (35) and there are (36) neutrons in an atom of this element.
7. This element has the maximum number of valence electrons in period 4. It is (37), symbol (38), atomic number (39), mass number (40) and it has (41) neutrons. Its family name is (42).
8. This element is a transition metal. It has (43) valence electrons and is found in period 5. It has a nuclear charge of +47. The element is (44), symbol (45) and atomic number (46). Its mass number is (47) and it has (48) neutrons in its nucleus.
9. This element has a mass number of 40 but it only has 2 valence electrons. It is not a transition metal. It is found in Group (49) and has electrons in (50) energy levels. The element is (51), symbol (52) and atomic number (53).
10. This element has the largest atoms of Period 4. The atomic number of this element is (54). The number of protons in one atom of this element is (55). An atom of this element has (56) valence electrons and has (57) completely filled quantum levels. The element is a member of the (58) family. The element is (59), symbol (60).
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